

Answer on Question#71535 – Chemistry – General chemistry

**Question:** 2,00 moles of H<sub>2</sub>O

**Solution:**

$$n(\text{H}_2\text{O}) = \frac{m(\text{H}_2\text{O})}{M(\text{H}_2\text{O})}$$

$$m(\text{H}_2\text{O}) = n(\text{H}_2\text{O}) \times M(\text{H}_2\text{O}) = 2.00 \text{ mol} \times 18.0 \frac{\text{g}}{\text{mol}} = 36.0 \text{ g}$$

**Answer:** 2.00 moles of H<sub>2</sub>O have mass 36.0 g.

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