

Answer on #71455, Chemistry / General Chemistry

Question : If the theoretical yield of a product reaction was 1.358g and the actual yield is 1.146 what is the percent yield of the product

Solution:

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

Actual yield is the amount produced in the experiment, and theoretical yield is the maximum amount possible based on calculations. So

$$\% \text{ yield} = (1.146 \text{ g}/1.358 \text{ g}) \times 100 = 84.39 \%$$

Answer: 84.39%

Answer provided by <https://www.AssignmentExpert.com>