## Answer on the question \#69468, Chemistry / Physical Chemistry

## Question:

Given the half redox reaction below, what is the coefficient of the electron in the balanced half reaction in a basic medium $\mathrm{CrO}^{\wedge}$ - $4+\mathrm{H} 2 \mathrm{O}$--> $\operatorname{Car}(\mathrm{OH}) 3$

## Answer:

Balanced half-reaction is:

$$
\mathrm{CrO}_{4}^{2-}+4 \mathrm{H}_{2} \mathrm{O}+3 e^{-} \rightarrow \mathrm{Cr}(\mathrm{OH})_{3}+5 \mathrm{OH}^{-}
$$

The coefficient of the electron is +3 .

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