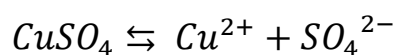


Answer on Question # 68839, Chemistry / General Chemistry

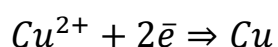
1) Two cells A and B are connected in series to a supply of electricity and 0.04 faraday of electricity are passed through each cell, each cell contains an aqueous solution of copper II tetraoxosulphate VII of the same concentration but cell A has copper electrodes and cell B has platinum electrode what happens at the anode in A and the cathode in B?

2) Also if cathode A as increased in mass by 1.25g, calculate the mass which would have been deposited by one Faraday of electricity at A.

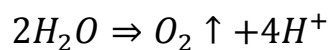
Solution:



1) Anode:



Cathode:



2)

$$m = \frac{1.25 \text{ g} \times 1 \text{ Faraday}}{0.04 \text{ Faraday}} = 31.25 \text{ (g)}$$

Answer: 31.25 g.