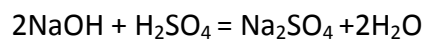


## Answer on Question #68027 - Chemistry -General Chemistry

how much .28 M NaOH is needed to neutralize 20.0 mL .56 M H<sub>2</sub>SO<sub>4</sub>?

**Solution:**



$$n(\text{NaOH}):n(\text{H}_2\text{SO}_4) = 2:1$$

$$n(\text{NaOH}) = 2 \times n(\text{H}_2\text{SO}_4)$$

$$c(\text{NaOH}) \times V(\text{NaOH}) = 2 \times c(\text{H}_2\text{SO}_4) \times V(\text{H}_2\text{SO}_4)$$

$$V(\text{NaOH}) = \frac{2 \times c(\text{H}_2\text{SO}_4) \times V(\text{H}_2\text{SO}_4)}{c(\text{NaOH})} = \frac{2 \times 0.56 \times 20.0}{0.28} = 80.0 \text{ ml}$$

**Answer:**

80.0 ml.



Answer provided by AssignmentExpert.com