Answer on Question #67389 - Chemistry -General Chemistry

How many grams of neon gas occupy 0.5 L at RTP? {NE: 20

Solution:

$$n(Ne) = \frac{V(Ne)}{V_m} = \frac{0.5}{22.711} = 0.022 \text{ mol}$$
$$m(Ne) = n(Ne) \times M(Ne) = 0.022 \times 20 = 0.44 \text{ g}$$

Answer: 0.44 grams.

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