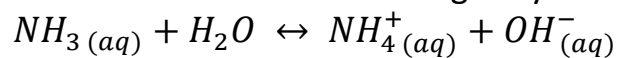


Answer on Question #67328, Chemistry / General Chemistry

For the following equation, indicate the original acid and base and the conjugate acid and base and indicate the reasoning for your answer.



Solution:

According to Brønsted-Lowry acid-base theory, the acid is H⁺-donor and the base is H⁺-acceptor. In our case, the donor is H₂O, and the acceptor is NH₃. For H₂O – acid the conjugated base must be OH⁻ and for NH₃ – base the conjugated acid must be NH₄⁺.

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