

How many moles are there in .028g of Fe?

Solution:

$$n = \frac{m}{M}$$
$$n(\text{Fe}) = \frac{m(\text{Fe})}{M(\text{Fe})} = \frac{0.028}{56} = 0.0005 \text{ (mol)}$$

Answer: $n(\text{Fe}) = 0.0005 \text{ mol}$