

how many atoms are there in 80.9 g of potassium oxide?

**Answer**

$$\nu = \frac{m}{Mr(K_2O)} = \frac{80.9}{94} = 0.86 \text{ mole}$$

$$\nu = \frac{N}{N_A};$$

$$N = \nu \cdot N_A = 0.86 \cdot 6.02 \cdot 10^{23} = 5.18 \cdot 10^{23}$$

$$N = 5.18 \cdot 10^{23} \cdot 3 = 1.55 \cdot 10^{24}$$

**Answer :**  $1.55 \cdot 10^{24}$  atoms

Answer provided by AssignmentExpert.com