Answer to the Question 66118

How many grams are required to make 2.8 L of 3.25% $AgNO_3$?

$$m = \rho \cdot V$$

$$\rho = 1000 g/L$$

$$m = 1000 g/L \cdot 2.8\, L\ = 2800 g$$

$$m(AgNO_3) = \frac{2800 g \cdot 3.25\%}{100\%} = 91 g$$