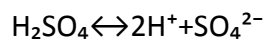


Answer on Question #64981, Chemistry / General Chemistry

Calculate the [OH<sup>-</sup>] for 0.0058 M H<sub>2</sub>SO<sub>4</sub>. Answer in unit M

**Answer**



$$[\text{H}^+] = 0,0058 * 2 = 0,0116 \text{ mol/L}$$

$$\text{pH} = -\lg[\text{H}^+] = 1,94$$

$$[\text{OH}^-] = 10^{-14} / 10^{-1,94} = 10^{-12,06} \text{ mol/L}$$

$$[\text{OH}^-] = 8,71 * 10^{-13} \text{ mol/L} = 87,1 * 10^{-3} \text{ M}$$

Answer provided by <https://www.AssignmentExpert.com>