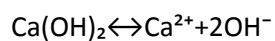


## Answer on Question #64980, Chemistry / General Chemistry

Calculate the  $[H_2O^+]$  for 0.018 M  $Ca(OH)_2$ . Answer in unit M

### Answer



$$[OH^-] = 0,018 \cdot 2 = 0,036 \text{ mol/L}$$

$$pOH = -\lg[OH^-] = 1,44$$

$$[H^+] = 10^{-14} / 10^{-1,44} = 10^{-12,56} \text{ mol/L}$$

$$[H^+] = 2,75 \cdot 10^{-13} \text{ mol/L} = 2,75 \cdot 10^{-13} \text{ M}$$

Answer provided by <https://www.AssignmentExpert.com>