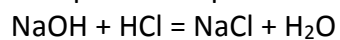


### Question #64033, Chemistry / General Chemistry

How do I calculate the amount of NaOH solution used, based on 0.999ml HCl in a titration experiment?

#### Answer:

At equivalence point:



$$n(\text{NaOH}) = n(\text{HCl})$$

$$c(\text{NaOH}) \times V(\text{NaOH}) = c(\text{HCl}) \times V(\text{HCl})$$

$$V(\text{NaOH}) = \frac{c(\text{HCl}) \times V(\text{HCl})}{c(\text{NaOH})}$$

c-concentration in M or mol/L

V-volume of solution in L