## Answer on Question #63897 Chemistry, General Chemistry

Provide the following utilizing your understanding of Avogadro's number:

-Mass, in grams, of  $3.0 \times 10^{-3}$  mol of ammonium phosphate

-Moles of chloride ions in 0.4 g of aluminum chloride.

## Answer:

- 1) n=m/M m=n·M M((NH<sub>4</sub>)<sub>3</sub>PO<sub>3</sub>) = 149 g/mol m=3.0 x  $10^{-3} \cdot 146 = 0.438$  g
- 2) n=m/M M(Cl) = 35.5 g/mol n (Cl') = 0.4/35.5 = 0.01 mol

Answer provided by www.AssignmentExpert.com