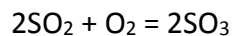
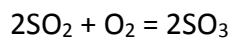


Answer on Question #63436 - Chemistry – General Chemistry

If 2.7 mol SO₂ is placed in a reaction vessel with an excess of O₂ gas, how many atoms of SO₃ are produced?



Solution.



- 1) 2 mol SO₂ – 2 mol SO₃
2.7 mol SO₂ – x mol SO₃

$$x = 2.7 \text{ mol SO}_3$$

- 2) $N(\text{SO}_3) = v(\text{SO}_3) \times N_a = 2.7 \times 6.02 \times 10^{23} = 16 \times 10^{23}$

Answer: $N(\text{SO}_3) = 16 \times 10^{23}$