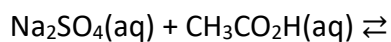


Answer on Question#63368 – Chemistry – General chemistry

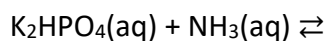
Question:

For each of the following reactions, write a balanced net ionic equation. Define whether the reaction goes nearly to completion or proceeds to only a small extent.



Write a balanced net ionic equation:

This reaction.



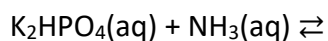
Write a balanced net ionic equation:

This reaction.

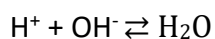
Answer:

The first reaction is impossible because ions that exist in the solution don't react each other. So, we cannot write a balanced net ionic equation.

There is the aqua solution of ammonia in the second reaction. Ammonia exist in hydrate form in this conditions: $\text{NH}_3 + \text{H}_2\text{O} \rightleftharpoons \text{NH}_3 \cdot \text{H}_2\text{O} \rightleftharpoons \text{NH}_4^+ + \text{OH}^-$



Write a balanced net ionic equation:



This reaction goes nearly to completion.