

Answer on Question #63274 - Chemistry - General Chemistry

Question:

Carbon dioxide, which is recognized as the major contributor to global warming as a "greenhouse gas," is formed when fossil fuels are combusted, as in electrical power plants fueled by coal, oil, or natural gas. One potential way to reduce the amount of CO₂ added to the atmosphere is to store it as a compressed gas in underground formations. Consider a 1000-megawatt coal-fired power plant that produces about 7×10^6 tons of CO₂ per year.

- 1) Assuming ideal gas behavior, 1.00 atm, and 17 °C, calculate the volume of CO₂ produced by this power plant.
- 2) If the CO₂ is stored underground as a liquid at 10 °C and 120 atm and a density of 1.2 g/cm³, what volume does it possess?
- 3) If it is stored underground as a gas at 33 °C and 90 atm, what volume does it occupy?

Solution:

- 1) The relation between pressure, volume, temperature and amount of ideal gas is expressed by the ideal gas equation:

$$P * V = (m / M) * R * T, \text{ where}$$

P – the pressure of gas (Pa),

V – volume of gas (m³),

m – mass of gas (g),

M – molar mass of gas (g/mol)

R – universal gas constant = 8.314 m³ Pa K⁻¹ mol⁻¹

T – absolute temperature (K).

Let's express volume from the equation above:

$$V = (m / M) * R * T / P$$

and bring to the standard units:

m = 7×10^6 tons = 7×10^{12} grams;

M (CO₂) = 12 + 16 * 2 = 44 g/mol;

T = 17 °C = (17 + 273) K = 290 K;

P = 1.00 atm = 101325 Pa.

Do the calculation:

$$V = (7 \times 10^{12} \text{ g} / 44 \text{ g/mol}) * 8.314 \text{ m}^3 \text{ Pa K}^{-1} \text{ mol}^{-1} * 290 \text{ K} / 101325 \text{ Pa} = 3.79 \times 10^9 \text{ m}^3.$$

- 2) Density is mass / volume. Therefore volume = mass / density.

Do the calculation:

$$V = 7 \times 10^{12} \text{ g} / 1.2 \text{ g/cm}^3 = 5.83 \times 10^{12} \text{ cm}^3 = 5.83 \times 10^6 \text{ m}^3.$$

3) Use the equation from part 1), but with another pressure and temperature:

$$P = 90 \text{ atm} = (90 * 101325) \text{ Pa} = 9.1 * 10^6 \text{ Pa};$$

$$T = 33 \text{ }^\circ\text{C} = (33+273) \text{ K} = 306 \text{ K}.$$

Do the calculation:

$$V = (7 * 10^{12} \text{ g} / 44 \text{ g/mol}) * 8.314 \text{ m}^3 \text{ Pa K}^{-1} \text{ mol}^{-1} * 306 \text{ K} / 9.1 * 10^6 \text{ Pa} = 4.45 * 10^7 \text{ m}^3.$$

Answer:

1) $3.79 * 10^9 \text{ m}^3$.

2) $5.83 * 10^6 \text{ m}^3$.

3) $4.45 * 10^7 \text{ m}^3$.

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