

Answer on Question #62906, Chemistry / General Chemistry

Find the mass of $6.3 \cdot 10^{22}$ molecules of SnCl_4

Solution:

$$6.3 \cdot 10^{22} \text{ molecules SnCl}_4 \times \frac{1 \text{ mol SnCl}_4}{6.02 \cdot 10^{23} \text{ molecules SnCl}_4} \times \frac{260.3 \text{ g SnCl}_4}{1 \text{ mol SnCl}_4} = 27.2 \text{ g SnCl}_4$$

Answer: 27.2 gSnCl₄

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