

**Answer on Question #59870 – Chemistry – General Chemistry**

**Question:**

How many moles are in 1.50g of CO<sub>2</sub>?

**Answer:**

$$n = \frac{m}{M}$$

where n – is a number of moles of a substance(mol), m – mass of a substance(g), M – molar mass of a substance (g/mol).

$$\text{Then } n = \frac{1.50}{44.01} = 0.034 \text{ moles}$$

**Answer:**      **0.034 moles** is the number of moles of **CO<sub>2</sub>**.