

**Answer on Question #59787 - Chemistry | General Chemistry**

calculate the number of moles in a 155.1 mg piece of aluminum foil

**Solution**

$$m(\text{Al}) = 155.1 \text{ (mg)} = 0.1551 \text{ (g)}$$

$$M(\text{Al}) = 26.98 \text{ (g/mol)}$$

$$n(\text{Al}) = \frac{m}{M} = \frac{0.1551 \text{ g}}{26.98 \text{ g/mol}} = 0.0057 \text{ (mol)} \approx 0.006 \text{ (mol)}$$

**Answer**

$$n(\text{Al}) = 0.006 \text{ (mol)}$$