Answer on Question #59503 - Chemistry - General Chemistry

How many moles of $CaBr_2$ are in 5.0 g of $CaBr_2$?

Solution:

M ($CaBr_2$) – 199.886 g/moles

$$N = \frac{m}{M}$$

$$N = \frac{5}{199.886} = 0.025 \ (moles)$$

Answer: In 5.0 g of $CaBr_2$ is 0.025 moles.