Question #58893, Chemistry / General Chemistry
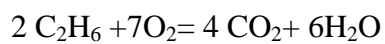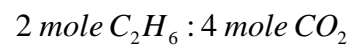
How many moles of CO2 are produced when 5.05 mol of ethane are burned in an excess of oxygen?

 Answer

2 C_2H_6 + 7O_2 = 4 CO_2 + 6H_2O

$$v = \frac{5.05 \cdot 4}{2} = 10.1 \, mole \, CO_2$$

$$2 \, mole \, C_2H_6 : 4 \, mole \, CO_2$$

$$Answer : v(CO_2) = 10.1 \, mole \, CO_2$$