

**Answer on Question #54250 – Chemistry – General Chemistry**

**Question:**

Please solve  $\frac{53.5\text{celcius}}{g} \times \frac{18.02g}{1.00\text{mole}} \times 1.00\text{Kcal}/1000.00\text{cal}$

**Answer:**

$$\begin{aligned} \frac{53.5\text{celcius}}{g} \times \frac{18.02g}{1.00\text{mole}} \times \frac{1.00\text{Kcal}}{1000.00\text{cal}} &= \frac{53.5\text{celcius}}{g} \times \frac{18.02g}{1.00\text{mole}} \times \frac{1000.00\text{cal}}{1000.00\text{cal}} \\ &= \mathbf{964.07 \frac{\text{celcius}}{\text{mole}}} \end{aligned}$$