

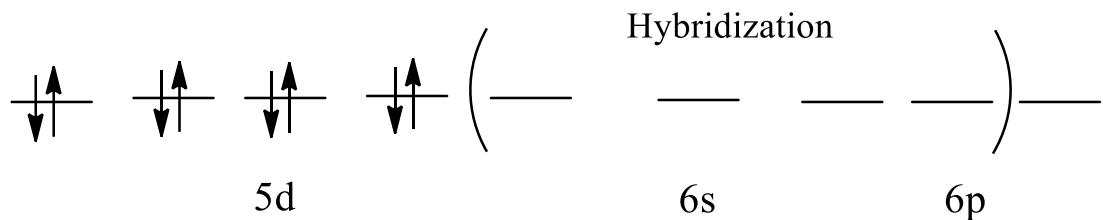
## Answer on Question #53714 – Chemistry – Inorganic Chemistry

### Question:

Find the hybridization of  $(PtCl_4)^{2-}$  by drawing the proper structure of it.

### Answer:

Having electronic configuration  $[Xe] 6s^0 4f^{14} 5d^8 6p^0$  Pt(II) has one s, two p and one d orbitals for hybridization.



Therefore, Pt has  $dsp^2$  hybridization, and the corresponding complex anion exhibits a square-planar geometry:

