Question #45233, Chemistry, Inorganic Chemistry

Molar mass of mixture of isotopes is equal to follow expression:

$$M = M_1\omega_1 + M_2\omega_2$$

Where M_1 and M_2 are molar mass of isotopes, ω_1 and ω_2 are mass ratio of isotopes in the mixture.

$$M_2 = \frac{M - M_1\omega_1}{\omega_2} = \frac{79.904 - 78.973 * 50.69}{49.31} = 81.1832$$ - is molar mass of Bromie-81