1. The reaction is:

$$10KClO_3(s) + 3P_4(s) \rightarrow 3P_4O_{10}(s) + 10KCl(s)$$

2. Convert to moles:

$$n_{\mathrm{KClO_3}} = \frac{m_{\mathrm{KClO_3}}}{M_{\mathrm{KClO_3}}} = \frac{52.9}{122.55} = 0.432\ (mole)$$

3. Multiply by mole ratio:

$$n_{P_4O_{10}} = n_{\mathrm{KClO_3}} \cdot \frac{3}{10} = 0.432\ \cdot 0.3 = 0.129\ (mole)$$

4. Convert to grams:

$$m_{P_4O_{10}} = n_{P_4O_{10}} \cdot M_{P_4O_{10}} = 0.129 \cdot 283.89 = 36.62\ (g)$$

Answer: 36.62 g.