I don't know, how to explain that SnO formation from its elements is thermodynamically favorable. I think it must be redox because that's how we explain thermodynamic favorable equations. But I don't know if I can use water molecules because it says "from its elements".

Solution: Such explanation should be based on the sign of the Gibbs free energy change for the reaction of SnO formation from elements in standard state (298 K, 1.013 bar):

$$2Sn_{(solid)} + O_{2(gas)} \longrightarrow 2SnO_{(solid)}$$

$\Delta G = \Delta H_r - T \cdot \Delta S_r$, where ΔG – Gibbs free energy change, kJ, if $\Delta G<0$, then this reaction is thermodynamically favorable; ΔH_r – reaction enthalpy change, kJ, $\Delta H_r = \Sigma(\Delta_f H^o_i \cdot n_i)_{products} - \Sigma(\Delta_f H^o_i \cdot n_i)_{reagents}$, where $\Delta_f H^o$ – standard enthalpy change of formation of substance, kJ/mol; n – substance amount, mol; T – absolute temperature, K; ΔS_r – reaction entropy change, kJ/K, $\Delta S_r = \Sigma(S^o_i \cdot n_i)_{products} - \Sigma(S^o_i \cdot n_i)_{reagents}$, where S^o – standard entropy of formation of substance, $kJ \cdot mol^{-1} \cdot K^{-1}$; n – substance amount, mol;

All the values of $\Delta_f H^o$ and S^o are available in handbooks.

We will not calculate the standard Gibbs free energy change of SnO formation, because it is too available in the handbooks, and it has value of –257.3 kJ/mol, as you can see, $\Delta G<0$, the reaction of SnO formation from its elements is thermodynamically favorable.

Answer: The standard Gibbs free energy change of SnO formation is negative, and then the reaction of SnO formation from its elements is thermodynamically favorable.